

11. The energy of micellization  $mX \rightleftharpoons X_m$

$$\Delta G = -\frac{RT}{m} \ln K = -\frac{RT}{m} \ln \frac{c_X}{m} \frac{1}{[c(1-x)]^m}$$

$$= -\frac{RT}{m} \left[ \ln \frac{c_X}{m} - m \ln c(1-x) \right]$$

at C.M.C.,  $x=0$ .

$$\Rightarrow \Delta G = RT \ln C.M.C.$$

$$\Delta S = -\frac{d\Delta G}{dT} = -RT \frac{d \ln C.M.C.}{dT} - R \ln C.M.C.$$

$$\Delta H = \Delta G + T\Delta S = RT \ln C.M.C. - RT^2 \frac{d \ln C.M.C.}{dT} - RT \ln C.M.C.$$

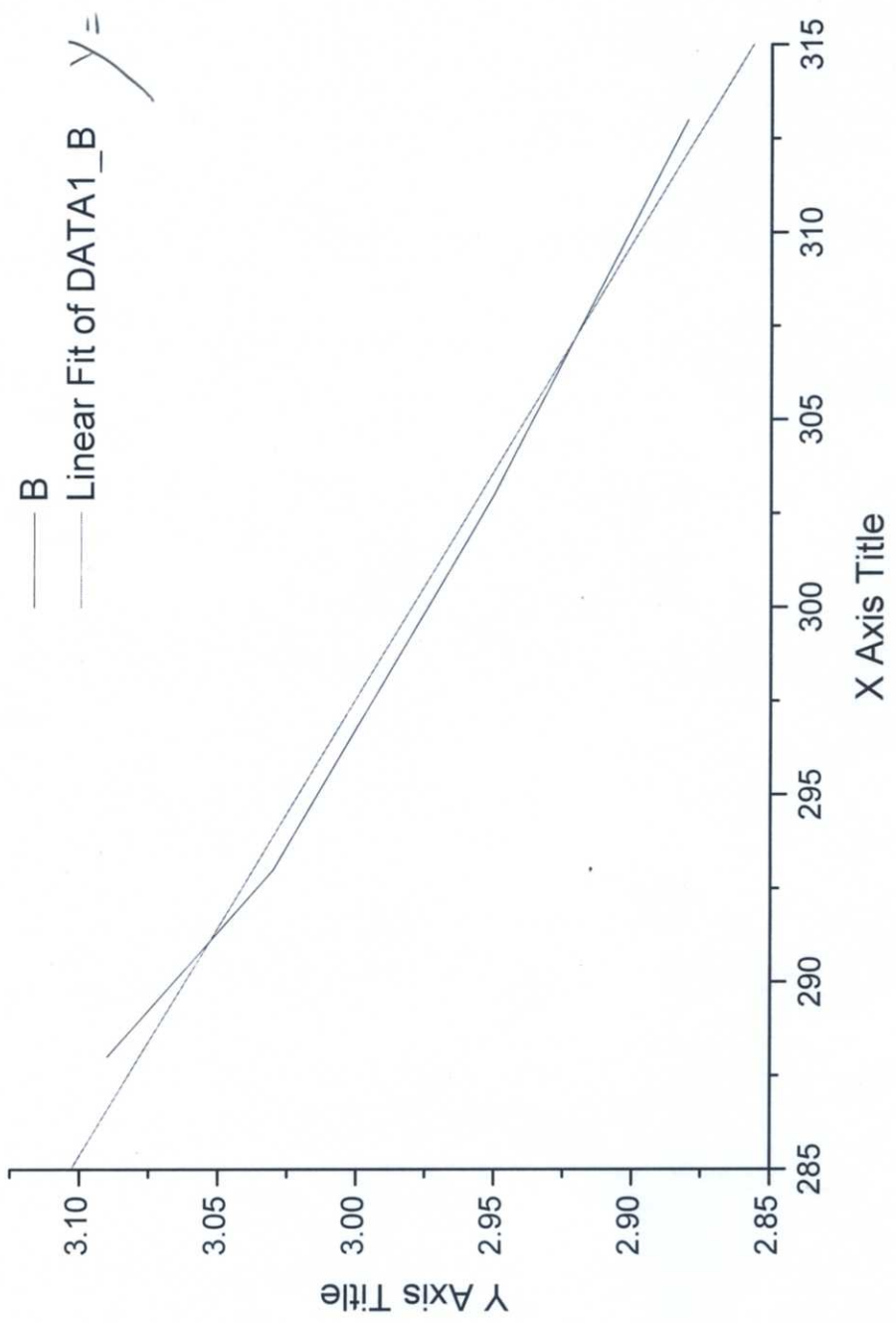
$$= -RT^2 \frac{d \ln C.M.C.}{dT}$$

$$\Rightarrow \Delta H = -RT^2 \frac{d \ln C.M.C.}{dT} \quad \leftarrow$$

T.(K)	288	293	298	303	313
C.M.C. (mmol/l)	1.9	20.8	19.9	19.1	12.8
$\ln C.M.C.$	3.09	3.03	2.99	2.95	2.88

$$\left( \Rightarrow \frac{d \ln C.M.C.}{dT} = -0.00822 \right)$$

$$\Delta H = -8.314 \times (300)^2 \times -0.00822 = 6150 \text{ J/mol} \quad \left( \frac{1}{2} \right)$$



$$y = 5.44465 \cdot 10^{-8} x + 0.00822$$

$$\Delta H = -RT^2 \frac{d \ln c.m.c}{dT} \quad \downarrow$$

$$-\frac{\Delta H}{RT^2} dT = d \ln c.m.c \quad \text{两边积分}$$

$$\Rightarrow \ln c.m.c = \frac{\Delta H}{R} \times \frac{1}{T} + C$$

$\ln c.m.c$	3.09	3.03	2.99	2.95	2.88
$\frac{1}{T}$	$3.47 \times 10^{-3}$	$3.41 \times 10^{-3}$	$3.36 \times 10^{-3}$	$3.3 \times 10^{-3}$	$3.19 \times 10^{-3}$

$$y = 0.5242 + 136.3x$$

$$\frac{dy}{dx} = \frac{\Delta H}{R} = 136.3$$

$$\Rightarrow \Delta H = 136.3 \times 8.314 = 6121 \text{ J/m.}$$

-2.



$$y = 0.524x + 236.3x$$

